

The Crystal and Molecular Structure of Tris(acetylacetonato)aquoytterbium(III) Hemibenzene¹

By EDWARD D. WATKINS, II, JAMES A. CUNNINGHAM, THEODORE PHILLIPS, II,
DONALD E. SANDS,² AND WILLIAM F. WAGNER

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The crystal structure of tris(acetylacetonato)aquoytterbium(III) hemibenzene, $\text{Yb}(\text{C}_5\text{H}_7\text{O}_2)_3(\text{H}_2\text{O}) \cdot 0.5\text{C}_6\text{H}_6$, has been determined from three-dimensional X-ray diffraction data. Two formula units are contained in a triclinic unit cell of space group $\text{P}\bar{1}$ and dimensions $a = 8.728 \pm 0.009 \text{ \AA}$, $b = 8.435 \pm 0.009 \text{ \AA}$, $c = 15.623 \pm 0.016 \text{ \AA}$, $\alpha = 95.95 \pm 0.02^\circ$, $\beta = 107.72 \pm 0.09^\circ$, and $\gamma = 96.18 \pm 0.09^\circ$. A total of 2625 intensities above background were collected by counter methods. The structure was refined by least-squares methods to a conventional R factor of 0.080; during the final cycle of refinement, corrections for anomalous dispersion were included. The ytterbium atom is surrounded by seven oxygen atoms at the vertices of a capped trigonal prism. These tris(acetylacetonato)aquoytterbium(III) molecules are linked in pairs by hydrogen bonds. The benzene molecule occupies a center of symmetry; all of the intermolecular contacts involving benzene carbon atoms and acetylacetonate carbon or oxygen atoms exceed 3.5 \AA .

Introduction

The preceding paper described the structure of $\text{Yb}(\text{C}_5\text{H}_7\text{O}_2)_3(\text{H}_2\text{O})$.³ In the course of preparing and studying the hydrates, it was observed that the crystals recovered from a solution of $\text{Yb}(\text{C}_5\text{H}_7\text{O}_2)_3 \cdot 3\text{H}_2\text{O}$ in benzene or acetone contained one molecule of the solvent for two molecules of $\text{Yb}(\text{C}_5\text{H}_7\text{O}_2)_3(\text{H}_2\text{O})$.⁴ The infrared spectra of these solvates differed from the spectrum of the monohydrate only in the appearance of additional bands characteristic of the solvent molecule. Interest in the possible effect of crystal packing on the molecular geometry prompted the determination of the structure of this benzene solvate.

Experimental Section

Crystals of $\text{Yb}(\text{C}_5\text{H}_7\text{O}_2)_3(\text{H}_2\text{O}) \cdot 0.5\text{C}_6\text{H}_6$ were prepared by Dr. Richardson.⁴ A crystal of dimensions $0.28 \times 0.23 \times 0.20 \text{ mm}$ was sealed in a glass capillary with the 0.28-mm dimension parallel to the capillary, and approximate unit cell dimensions and the diffraction symmetry were obtained from oscillation, Weissenberg, and precession photographs. Accurate unit cell dimensions and the intensity data were measured on a General Electric X-ray spectrometer equipped with a single-crystal orienter and a scintillation counter. The crystals are triclinic with dimensions (based on $\text{Mo K}\alpha = 0.7107 \text{ \AA}$) $a = 8.728 \pm 0.009 \text{ \AA}$, $b = 8.435 \pm 0.009 \text{ \AA}$, $c = 15.623 \pm 0.016 \text{ \AA}$, $\alpha = 95.95 \pm 0.02^\circ$, $\beta = 107.72 \pm 0.09^\circ$, and $\gamma = 96.18 \pm 0.09^\circ$. The long dimension of the crystal was parallel to a . With two molecules of $\text{Yb}(\text{C}_5\text{H}_7\text{O}_2)_3(\text{H}_2\text{O}) \cdot 0.5\text{C}_6\text{H}_6$ per unit cell the calculated density is 1.62 g cm^{-3} .

The intensities were measured manually by the stationary-crystal, stationary-counter technique using $\text{Mo K}\alpha$ radiation and 40-sec counts. The crystal was oriented about a^* , and the settings were calculated for the $\text{K}\alpha_1$ line ($\lambda 0.70930 \text{ \AA}$). Conversion of peak heights to integrated intensities were based upon several intensities measured by a manual ω scan. Background corrections were obtained by computer interpolation of a plot of intensity *vs.* angle, constructed from measurements made with the

crystal out of reflecting position. No corrections were made for absorption ($\mu = 45.9 \text{ cm}^{-1}$) or extinction. The estimated maximum variation in the ratio of any two intensities due to absorption is 10.5%.

It was initially planned to measure all reflections for which 2θ was less than 50° . However, problems in electronic stability which arose during the measurements indicated the advisability of completing the data collection as soon as possible, so the maximum 2θ was decreased to 35° . A total of 2963 reflections were measured, and these included about 60% of the accessible reflections between 35 and 50° . Of the measured reflections, 2625 were at least 7% above background and were classified as observed. The variation in the reflected beam intensity was 32%.

Structure Determination and Refinement

Space group $\text{P}\bar{1}$ was assumed and was ultimately confirmed by the structure. Trial ytterbium positions were deduced from the three-dimensional Patterson function. Further interpretation of the Patterson function was not feasible. The ytterbium atoms occupy the $\text{P}\bar{1}$ general positions $2(i): \pm(x, y, z)$ with $x = 0.012$, $y = 0.190$, $z = 0.165$. The ytterbium contributions to the structure factors were calculated, and a difference Fourier synthesis was computed using the 2539 reflections for which $|F_o| \geq 0.40|F_c|$. This difference Fourier map revealed the positions of the seven oxygen atoms. Two additional difference Fourier maps were required before all of the carbon atoms were located.

Least-squares refinements were carried out using a local IBM 360 adaptation of the program of Busing, Martin, and Levy.⁵ The refinements included one scale factor, the positional parameters for one ytterbium atom, seven oxygen atoms, and eighteen carbon atoms, anisotropic temperature factors for the ytterbium atom, and isotropic temperature factors for the oxygen and carbon atoms. The atomic form factors used were those of Cromer, Larson, and Waber⁶ for neutral ytterbium and from the compilation of Ibers⁷ for oxy-

(1) Work supported by U. S. Atomic Energy Commission Contract No. AT(40-1)2124. We also thank the University of Kentucky Computing Center.

(2) Author to whom correspondence should be addressed.

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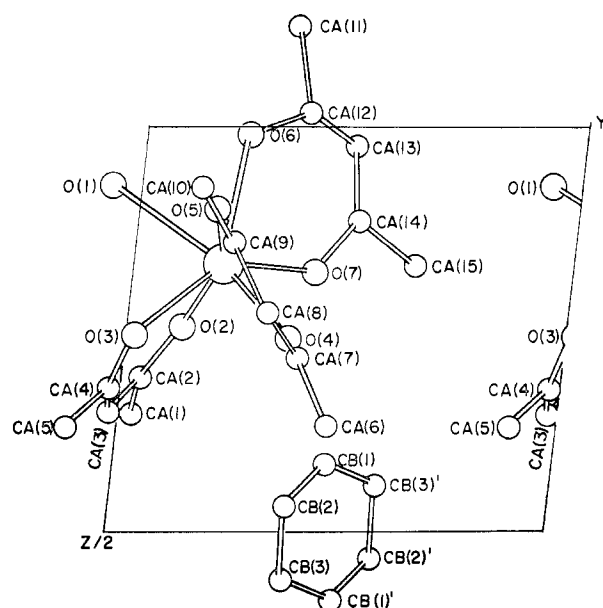


Figure 1.—Projection of a molecule along [100]. The primed and unprimed atoms are related by a center of symmetry.

TABLE III
BOND LENGTHS AND ANGLES, CHEMICALLY
EQUIVALENT DISTANCES BEING GROUPED TOGETHER

Lengths, Å ^a			
Yb-O(1)	2.347(11)	CA(4)-CA(3)	1.44(2)
Yb-O(2)	2.204(12)	CA(7)-CA(8)	1.37(3)
Yb-O(3)	2.278(13)	CA(9)-CA(8)	1.49(2)
Yb-O(4)	2.231(12)	CA(12)-CA(13)	1.35(2)
Yb-O(5)	2.252(11)	CA(14)-CA(13)	1.37(2)
Yb-O(6)	2.258(12)	CA(2)-CA(1)	1.53(3)
Yb-O(7)	2.250(13)	CA(4)-CA(5)	1.58(3)
O(2)-CA(2)	1.24(2)	CA(7)-CA(6)	1.50(3)
O(3)-CA(4)	1.28(2)	CA(9)-CA(10)	1.58(3)
O(4)-CA(7)	1.25(2)	CA(12)-CA(11)	1.55(2)
O(5)-CA(9)	1.32(2)	CA(14)-CA(15)	1.56(3)
O(6)-CA(12)	1.34(2)	CB(1)-CB(2)	1.33(3)
O(7)-CA(14)	1.28(2)	CB(2)-CB(3)	1.39(3)
CA(2)-CA(3)	1.41(2)	CB(3)-CB(1)'	1.54(3)

Angles, degrees			
O(2)-Yb-O(3)	74.8(4)	O(4)-CA(7)-CA(6)	120.4(21)
O(4)-Yb-O(5)	77.2(4)	O(5)-CA(9)-CA(10)	111.0(15)
O(6)-Yb-O(7)	74.5(4)	O(6)-CA(12)-CA(11)	112.6(17)
		O(7)-CA(14)-CA(15)	113.5(17)
Yb-O(2)-CA(2)	138.2(13)	CA(1)-CA(2)-CA(3)	115.1(18)
Yb-O(3)-CA(4)	133.9(14)	CA(5)-CA(4)-CA(3)	119.0(19)
Yb-O(4)-CA(7)	135.2(14)	CA(6)-CA(7)-CA(8)	115.6(21)
Yb-O(5)-CA(9)	128.5(11)	CA(10)-CA(9)-CA(8)	124.0(17)
Yb-O(6)-CA(12)	134.7(12)	CA(11)-CA(12)-CA(13)	122.4(19)
Yb-O(7)-CA(14)	136.2(13)	CA(15)-CA(14)-CA(13)	120.8(18)
O(2)-CA(2)-CA(3)	125.4(19)	CA(2)-CA(3)-CA(4)	120.6(18)
O(3)-CA(4)-CA(3)	125.1(20)	CA(7)-CA(8)-CA(9)	123.5(20)
O(4)-CA(7)-CA(8)	124.0(21)	CA(12)-CA(13)-CA(14)	123.8(20)
O(5)-CA(9)-CA(8)	125.0(17)	CB(3)' ^a -CB(1)-CB(2)	122.1(27)
O(6)-CA(12)-CA(13)	125.0(18)	CB(1)-CB(2)-CB(3)	117.0(30)
O(7)-CA(14)-CA(13)	125.7(19)	CB(2)-CB(3)-CB(1)'	120.7(26)
O(2)-CA(2)-CA(1)	119.4(19)		
O(3)-CA(4)-CA(5)	115.9(19)		

^a Primed and unprimed atoms are related by a center of symmetry.

O(1)-O(5)' = O(1)'-O(5) = 2.66 Å, where the primes indicate atoms related to the unprimed atoms by the center of symmetry at 0, 0, 0. The distance between water molecules is O(1)-O(1)' = 3.27 Å. The closest contact between pairs of molecules is 3.60 Å between O(2) and the CA(10) methyl group in the molecule generated by a lattice translation along *a*.

The coordination polyhedron formed by the seven oxygen atoms around an ytterbium atom is a capped trigonal prism (Figure 3), with O(3) capping the face O(1)-O(2)-O(4)-O(5). The dimensions of the co-

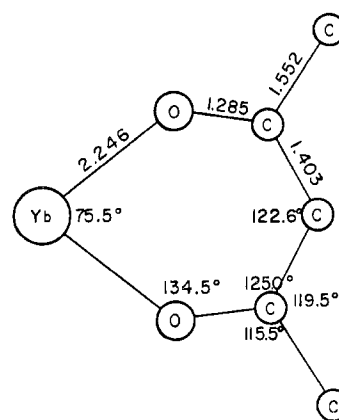


Figure 2.—Average dimensions of a ring in $\text{Yb}(\text{C}_6\text{H}_7\text{O}_2)_3(\text{H}_2\text{O}) \cdot 0.5\text{C}_6\text{H}_6$.

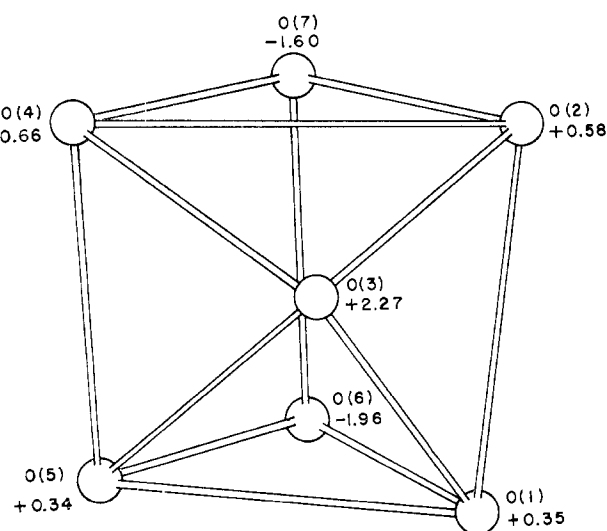


Figure 3.—Coordination polyhedron formed by the seven oxygen atoms bonded to the ytterbium atom. The distance of each atom from the central plane is given. The view is along a vector 30° from Yb-O(6) in the Yb-O(6)-O(7) plane.

TABLE IV
DIMENSIONS OF COORDINATION POLYHEDRON
LENGTHS, Å^a

O(1)-O(2)	3.09	O(4)-O(7)	2.86
O(1)-O(3)	2.86	O(7)-O(2)	2.86
O(1)-O(5)	2.98	O(7)-O(6)	2.73
O(1)-O(6)	2.77	O(2)-O(3)	2.72
O(4)-O(2)	3.51	O(3)-O(5)	2.92
O(4)-O(3)	2.83	O(5)-O(6)	2.84
O(4)-O(5)	2.80	O(2)-O(6)	3.85
O(4)-O(6)	3.93		

ANGLES, DEGREES ^b			
O(4)-O(2)-O(3)	52.0	O(1)-O(2)-O(3)	58.6
O(2)-O(3)-O(4)	78.5	O(2)-O(3)-O(1)	67.2
O(2)-O(4)-O(3)	49.4	O(2)-O(1)-O(3)	54.2
O(4)-O(3)-O(5)	58.2	O(1)-O(3)-O(5)	62.0
O(3)-O(5)-O(4)	59.2	O(3)-O(5)-O(1)	58.0
O(3)-O(4)-O(5)	62.6	O(3)-O(1)-O(5)	59.9
O(4)-O(5)-O(6)	88.5	O(1)-O(5)-O(6)	56.7
O(5)-O(6)-O(4)	45.3	O(5)-O(6)-O(1)	64.3
O(5)-O(4)-O(6)	46.2	O(5)-O(1)-O(6)	59.0
O(4)-O(7)-O(6)	89.4	O(1)-O(2)-O(6)	45.3
O(7)-O(6)-O(4)	43.9	O(2)-O(6)-O(1)	52.7
O(7)-O(4)-O(6)	46.7	O(2)-O(1)-O(6)	82.0
O(4)-O(7)-O(2)	75.8	O(2)-O(6)-O(7)	47.8
O(7)-O(2)-O(4)	52.0	O(6)-O(7)-O(2)	87.1
O(7)-O(4)-O(2)	52.1	O(6)-O(2)-O(7)	45.1

a) STANDARD DEVIATION = 0.02 Å.
b) STANDARD DEVIATIONS ARE 0.3° TO 0.5°.

ordination polyhedron are given in Table IV. This polyhedron is identical, within experimental error, with the polyhedron found about Yb(1) in the un-

solvated $\text{Yb}(\text{C}_6\text{H}_7\text{O}_2)_3(\text{H}_2\text{O})$ structure,³ and confidence in the reality of such features of the polyhedron as the deviations of the ytterbium atom from the mean planes of the ligands may be obtained by comparing the two structures.

The benzene molecule is centered on the center of symmetry at $0, \frac{1}{2}, \frac{1}{2}$. The mean deviation of the carbon atoms from the mean plane of the benzene ring is 0.019 \AA , and the sum of the interior angles of the ring is 719.6° . Discrepancies in the carbon-carbon bond lengths (Table III) of several standard deviations must be interpreted as indicative of systematic errors in the data, such as uncorrected adsorption, rather

than as implying any actual distortion of the molecule. The shortest distance between a benzene carbon atom and another carbon or oxygen atom is 3.53 \AA for the CB(1)-O(4) contact.

The volume occupied by two tris(acetylacetonato)-aquoytterbium(III) molecules in the unsolvated structure³ is 931.8 \AA^3 . The molecular volume of benzene, based on the density of the liquid at 25° , is 147.6 \AA^3 . The sum of these volumes is 1079.4 \AA^3 , as compared with a unit cell volume of 1077.9 \AA^3 in the crystal structure of the benzene solvate. These figures suggest that the efficiency of packing of the molecules is about the same for the solvated and unsolvated structures.

CONTRIBUTION FROM THE LOS ALAMOS SCIENTIFIC LABORATORY,
UNIVERSITY OF CALIFORNIA, LOS ALAMOS, NEW MEXICO 87544

The Crystal Structure of Ammonium Hexafluorocerate(IV), $(\text{NH}_4)_2\text{CeF}_6$ ¹

By ROBERT R. RYAN, ALLEN C. LARSON, AND F. H. KRUSE

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The crystal structure of ammonium hexafluorocerate(IV), $(\text{NH}_4)_2\text{CeF}_6$, has been determined from counter measurements of 561 reflections and refined by full-matrix least squares to an R factor of 0.032. The space group is Pbcn (no. 60); $D_x = 3.05$ and $D_m = 3.05 \pm 0.02 \text{ g/cm}^3$. The unit cell has dimensions $a = 7.026$ (3), $b = 12.098$ (6), and $c = 7.439$ (8) \AA ; $Z = 4$. Of the 561 reflections measured, 389 were judged to be observed. The structure consists of infinite chains of eightfold coordinated cerium atoms parallel to the c axis, joined by NH_4^+ ions. The fluorine configuration about the Ce can be roughly described as a square antiprism. The average Ce-F distance is 2.26 \AA and the average F-F distance for nearest neighbors within an antiprism is 2.77 \AA . Hydrogen bonding causes this particular structure to be stabilized over the trigonal-dodecahedral arrangement found in the similar Rb_2UF_6 and K_2ZrF_6 compounds.

Introduction

The preparation of ammonium hexafluorocerate(IV) was first reported by Besse and Capestan,² who indexed the powder pattern on the basis of a cubic cell with $a' = 10.52 \text{ \AA}$ and claimed a measured density of 3.27 g/cm^3 . For $Z = 8$, $D_x = 3.31 \text{ g/cm}^3$ for this cell. Penneman³ of this laboratory noticed that these values for the volume and Z give a very low value for the fluorine volume, as compared to similar actinide fluoride compounds. Further, there is no evidence for cubic phases of any other 2:1 NH_4F -actinide(IV) fluoride complexes. He was therefore prompted to prepare this material. Optical examination showed single crystals of $(\text{NH}_4)_2\text{CeF}_6$ to be biaxial and thus this compound must have orthorhombic or lower symmetry. Subsequent examination of precession films showed the symmetry to be orthorhombic with cell dimensions which are nearly those which would permit indexing of powder data on the basis of the cubic cell of Besse and Capestan,² i.e., $a'/a = \frac{3}{2}$, $a'/b = \sqrt{3}/2$, and $a'/c = \sqrt{2}$.

The single-crystal X-ray investigation of this compound was undertaken for three reasons: (1) No

compound of known structure could be found with which it is isostructural. (2) Comparison of the powder patterns indicated it is isostructural with the α phase of $(\text{NH}_4)_2\text{UF}_6$ for which no single crystals were then available. (Subsequent comparisons of single-crystal data, when they became available, confirmed that the two compounds were indeed isostructural. This is of interest because of the existence of four closely related polymorphs for $(\text{NH}_4)_2\text{UF}_6$ over a narrow temperature range⁴ ($\sim 70^\circ$), and the elucidation of these structures should provide information concerning the energy relationships between some of the possible coordination polyhedra for uranium.) (3) It is also of interest that after extensive investigation the only phase to be found for the Ce compound is the α phase. It is hoped that the knowledge of this structure plus the structure of the γ - $(\text{NH}_4)_2\text{UF}_6$ (room-temperature stable phase) compound currently being worked on will provide some information concerning these questions.

Experimental Section

Crystals were prepared by cooling a saturated solution of $(\text{NH}_4)_2\text{CeF}_6$ in a 15% NH_4F - H_2O solution. A needlelike crystal was selected and cut to a length suitable for examination by X-

(1) This work was sponsored by the U. S. Atomic Energy Commission.

(2) J. P. Besse and M. Capestan, *Bull. Soc. Chim. France*, 577 (1966).

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